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SPECIFICATION

PRECIOUS METAL-BASED AMORPHOUS ALLOY

TECHNICAL FIELD TO WHICH THE INVENTION PERTAINS

5 The present invention relates to a precious metal-based amorphous alloy used as a material for accessories or medical devices. Specifically, the present invention relates to a precious metal-based amorphous alloy rich in precious metal components and free of nickel which may have an influence on
10 the human body.

BACKGROUND ART

Precious metals such as platinum and palladium have been used for medical devices such as dental instruments and
15 catheters in addition to accessories such as rings, necklaces and pendants. Each of the materials used for these applications is required to have a higher hardness because the material needs to be prevented from scoring which is caused by the friction in use. A pure precious metal, which
20 is soft and vulnerable, is generally alloyed with a small amount of other metal elements when the precious metal is used as a material for the accessories and the medical devices. However, thus prepared precious metal alloys do not always have a fully satisfying property in terms of hardness.

25 A crystal structure of an amorphous alloy which is also referred to as a super-cooled metal or a glass metal is different from that of a general metal material, and this

amorphous alloy is a material having a random atomic arrangement throughout the wide range. This structure provides some features that defects which would otherwise exist in its crystal structure (grain boundaries, dislocations) can not be seen, that its physical characteristics such as strength show specific tendencies, and that particularly its hardness becomes extremely high. This amorphous alloy is manufactured by super-quenching the liquid state alloy, so that the cooling rate in this case is required to be at an adequate level for inhibiting the production of crystal nuclei and their growth (a critical cooling rate) (for example, a critical cooling rate for a precious metal alloy is approximately 10^2 to 10^4 °C/sec. and critical cooling rates for other alloys are approximately 10^5 to 10^6 °C/sec.). Such a limitation on the cooling rate has so far resulted in a restriction of a size of the amorphous alloy which can be manufactured, that is, only some types of materials including foil-like, needle-like, and flake-like materials can be manufactured, so that it has been difficult to use these alloys industrially.

However, with respect to an alloy metal having a predetermined composition, it has been recently found out that its material structure can be made into an amorphous state even at a relatively low cooling rate. This results in the manufacture of a bulky (ingot-like) and thick amorphous alloy which is larger than the size of the hitherto known amorphous alloy such as a foil type material. As an alloy

composition having such an ability of forming the amorphous state, various kinds of alloys have already been known. And applications of the amorphous alloys to the above described materials for accessories or medical devices, for example, are now under investigation.

As an example of studying an amorphous alloy which contains a precious metal, for example, a Pd-Ni-P based amorphous alloy (in atomic %, Pd 40%, Ni 40%, and P 20%) is described in Japanese Patent Laid-Open No. 59-35417 as one of the transition metal-semi metal based amorphous alloys. Using the precious metal alloy having this composition, it has been demonstrated that the amorphous alloy about 5 mm in thickness can be manufactured even by the metal mold casting. In addition, Japanese Patent Laid-Open No. 9-195017 describes a Pt-Pd-Cu-Si based amorphous alloy (in atomic %, Pt+Pd: 65 to 80%, Cu: 0 to 15%, and Si: 10 to 20%) and discloses that the precious metal alloy having this composition can also be made into a bulk of 100 mm in length and 1 mm in diameter.

However, these conventional amorphous alloys containing the precious metals are insufficient when considering their applications to the materials used for the accessories and the medical devices as described above. For example, the accessory is frequently desired to have an asset value as its aspect, and this asset value is commonly supposed to become greater in proportion to an amount of the precious metal contained in the accessory. Many of the conventional amorphous alloys contain less precious metals, so that in this

respect it can hardly be said that these amorphous alloys are suitable for the materials used for the accessories.

In addition, many of the above described conventional precious metal-based amorphous alloys contain nickel as their components, but nickel is an element whose influence on the human body such as an metal allergy and carcinogenesis is worried. Therefore, it can be considered that these conventional amorphous alloys are not favorable to be used for substances which are in contact with the human skin continuously such as accessories and for substances which are in contact with the internal tissue of the human body of the human such as medical devices.

The present invention has developed under the background as described above, and an object of the present invention is to provide an amorphous alloy which is rich in precious metals and is completely free of nickel provided that a bulk having an amorphous structure can be formed even when the alloy is solidified at a relatively low cooling rate.

20 DISCLOSURE OF THE INVENTION

The inventors have intensively made an effort to develop a precious metal-based amorphous alloy by which the above described problems can be solved. Specifically, the inventors have achieved the present invention as a result of selecting platinum as the precious metal which constitutes a principal component of the alloy, platinum being the most common material for accessories, to allow platinum to be

contained at a level of 50% or more of the alloy, as well as selecting Cu and P as additional elements which have the ability to form the amorphous structure, and variously changing the concentrations of these elements to investigate the respective structures of the alloys.

A first precious metal-based amorphous alloy according to the present application is a precious metal-based amorphous alloy with a Pt-Cu-P based structure comprising $50\% \leq \text{Pt} \leq 70\%$ by atom, $5\% \leq \text{Cu} \leq 35\%$ by atom, and $15\% \leq \text{P} \leq 25\%$ by atom.

A second precious metal-based amorphous alloy according to the present application is a precious metal-based amorphous alloy with a Pt-Pd-Cu-P based structure comprising $5\% \leq \text{Pt} \leq 70\%$ by atom, $5\% \leq \text{Pd} \leq 50\%$, $5\% \leq \text{Cu} \leq 50\%$ by atom, and $5\% \leq \text{P} \leq 30\%$ by atom.

An exact mechanism of forming the amorphous structure with respect to these two kinds of precious metal alloys according to the present invention is not completely revealed, but it is supposed that copper and phosphorus both of which are additional elements have some effects of raising the crystallization temperature of the alloy and of expanding the temperature range of a super-cooled liquid (a difference between the crystallization temperature and the glass transition temperature) of the above described alloy, so that the ability of forming the amorphous structure is improved.

In addition, the precious metal-based alloy with the Pt-Cu-P based structure and the precious metal-based alloy

with the Pt-Pd-Cu-P based structure according to the present invention can be made into amorphous states even when their cooling rates are relatively low by, as for the Pt-Cu-P based structure, defining a range of copper concentration as $5\% \leq \text{Cu} \leq 35\%$ and a range of phosphorus concentration as $15\% \leq \text{P} \leq 25\%$ provided that a concentration of platinum is 50% or more and 75% or less and by, as for the Pt-Pd-Cu-P based structure, defining a range of copper concentration as $5\% \leq \text{Cu} \leq 50\%$ and a range of phosphorus concentration as $5\% \leq \text{P} \leq 30\%$ provided that a concentration of platinum is 5% or more and 70% or less and a concentration of palladium is 5% or more and 50% or less. That is, if at least one of these constituents becomes outside of the above described range, the alloy is crystallized and its amorphous structure can not be obtained.

Although the precious metal-based amorphous alloys according to the present invention can be made into a bulky material even when the alloy is cooled at a relatively low cooling rates such as 10^{20}C/sec. or less, the alloy has a preferable cooling rate in order to more reliably obtain its amorphous structure. For example, in particular, a cooling rate for the Pt-Cu-P based structure is preferably from 10^{-1} to 10^{20}C/sec. , and a cooling rate for the Pt-Pd-Cu-P structure is preferably from 10^1 to 10^{20}C/sec. The amorphous alloy which has been cooled at this cooling rate is the precious metal-based alloy which has been completely made into its amorphous state because the cooling rate during its solidification is defined within an appropriate range. The

amorphous alloy according to the present invention which is thus completely made into its amorphous state has an extremely high hardness and is suitable for a material used for accessories or medical devices.

5 The precious metal-based amorphous alloy according to the present invention can contain up to 75% or 70% of platinum. Therefore, if the alloy is used for the accessories, an amount of the platinum contained therein can be expected to provide the accessories with the asset values. In addition, the
10 precious metal-based amorphous alloy according to the present invention is completely free of nickel as is evident from its composition, so that the alloy is supposed to have no effects on the human body which would otherwise cause metal allergy or carcinogenesis. In this respect, it also becomes possible
15 to use the alloy for accessories and medical devices.

 In addition, when each of the Pt-Cu-P based amorphous alloy and the Pt-Pd-Cu-P based amorphous alloy according to the present invention is made into its final product shape through casting, a surface of the alloy after being solidified
20 becomes smooth, so that the alloy can be used as a product as it is. In addition, the plastic workability of the amorphous alloy according to the present invention depends on its composition, but in the case where the alloy needs to be subjected to the strong working, its workability can be
25 retained by heating the alloy to a certain temperature between its glass transition temperature and its crystallization temperature (a supercooling liquid temperature range) for

performing the working. This results from a superplasticity phenomenon which is caused by an abrupt reduction in a viscosity of the amorphous alloy due to the heating.

As a method for manufacturing the precious metal-based
5 amorphous alloy according to the present invention, the alloy
can be manufactured by mixing each metal and phosphorus within
a predetermined range of the composition and by quenching the
molten metal with this composition before solidifying the
molten metal. When raw materials are mixed with and melted
10 into each other, it is preferable to use powdery raw materials
in order to promote the melting process. In this case, Cu
which is in a pure metal state can be added, but Cu which is
in a state of a copper-phosphide compound (Cu_3P and the like)
can be added in order to make fine adjustments of the
15 phosphorus concentration. Further, when these metals are
allowed to be melted, it is preferable to add borax in order
to prevent the alloy from oxidation. Although there is no
particular problem about a method which is to be performed
for quenching the alloy after the melting, a method for
20 rapidly casting the alloy into a copper mold after the alloy
is melted in a crucible being made of quartz for example or
a method for dipping a crucible in water is given as an example
of the methods being capable of cooling the alloy at a cooling
rate which is within a favorable range of temperature for each
25 of the above described alloy structure (10^{-1} to 10^{20} C/sec. for
the Pt-Cu-P based structure and 10^1 to 10^{20} C/sec. for the
Pt-Pd-Cu-P based structure).

BRIEF DESCRIPTION OF THE DRAWING

FIG. 1 shows a DSC curve of a specimen No. 7 (Pt: 60 at%, Cu: 20 at%, P: 20 at%).

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MODE FOR CARRYING OUT THE INVENTION

Preferable embodiments according to the present invention will now be described below with reference to the drawings. In the present embodiment, two kinds of precious metal-based amorphous alloys, one of which being Pt-Cu-P based structure and the other of which being Pt-Pd-Cu-P based structure, were manufactured, and a degree of amorphous state (hereinafter referred to as a vitrification degree) and a hardness of each of the alloys was measured to determine a composition range of the alloy having an amorphous structure.

EXAMPLE 1

In this example, Pt-Cu-P based amorphous alloys having different compositions were manufactured. After platinum powder, powdery red phosphorus, and small bulky copper phosphide (Cu_3P) were weighed so that a total amount of these materials became 100 g in order to obtain a composition described in Table 1 and mixed with each other, 5 g of borax were further added to the mixture, then the mixture was placed in a one-side sealed-off silica tube having an inner diameter of 20 mm to heat it within an electric furnace in an atmosphere of argon and allow the materials to be melted. The melting

temperature was determined to be 1100°C, and after the materials were melted at this temperature, an argon gas was blown into the molten metals and bubbling was allowed to be generated for one minute in order to stir and degas of the molten metals. Next, this molten metal was cast into a copper mold whose recess was in a ring shape (20 mm in outer diameter, 15 mm in inner diameter, and 50 mm in depth), and quenched and solidified to manufacture a ring shaped amorphous alloy.

With respect to each of the amorphous alloys thus manufactured, after the alloy was cut into a predetermined dimension, a differential thermal analysis was conducted, then a vitrification degree of each alloy was investigated while measuring its glass transition temperature and crystallization temperature. In this case, the differential thermal analysis was conducted by heating this alloy assuming that the weight of each amorphous alloy was constant within a range of 100 mg \pm 10 mg, and the vitrification degree was determined from a height of an exothermic peak which may appear during the crystallization. For example, a specimen No. 7 (Pt: 60 at%, Cu: 20 at%, P: 20 at%) of FIG. 1 shows that its glass transition temperature is 238.5°C and its crystallization temperature is 286.0°C. In addition, after this determination of the vitrification degree was performed, a Vickers hardness of each alloy described above was measured. Both results of measuring the vitrification degree and the Vickers hardness with respect to each alloy described above are shown in Table 1.

[illegible]

◎: Completely vitrified
 ○: Almost vitrified
 ×: Crystallization

As a result of this, an amorphous alloy having a composition within a range recited in claim 1 had a good vitrification degree and could be easily made into an amorphous structure, in addition, the alloy whose hardness is higher than a hardness of a platinum pure metal or a platinum alloy could be obtained. Every alloy was excellent in its gloss.

Also, the specimen No. 7 had a density of 15.39 g/cc. Investigating the mechanical characteristics of this specimen No. 7 which was molded into a ring shape having an outer diameter of 20.0 mm, an inner diameter of 16.0 mm, and a width of 3.0 mm, its compressive strength was 56 kg/cm². This alloy may have inscriptions thereon and its hardness and compressive strength are both higher than the platinum alloy, so that this alloy is considered to be suitable for the materials used for accessories.

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EXAMPLE 2

In this example, Pt-Pd-Cu-P based amorphous alloys which had different compositions described in Table 2 were manufactured. As is the case with Example 1, after platinum powder, powdery palladium, powdery red phosphorus, and small bulky copper phosphide (Cu₃P) were weighed so that a total amount of these materials became 100 g in order to obtain a composition described in Table 2 and mixed with each other, 5 g of borax were further added to the mixture, then the mixture was placed in a one-side sealed-off silica tube having an inner diameter of 20 mm to heat it within an electric furnace at 1100°C in an atmosphere of argon and allowed the materials to be melted. An argon gas was blown into the molten metals and bubbling was allowed to be generated for one minute. Next, this molten metal was dipped in water together with the silica tube, and quenched and solidified to manufacture a rod-like amorphous alloy.

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After each of these amorphous alloys was cut into a predetermined dimension, a differential thermal analysis was conducted, then a vitrification degree of each alloy was investigated while measuring its glass transition

5 temperature and crystallization temperature. Analytical conditions were the same as in the case of Example 1. Both results of measuring the vitrification degree and the Vickers hardness with respect to each alloy manufactured in this example are both shown in Table 2.

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[Table 2]

Specimen No.	Element concentration (at%)				Degree of vitrification (Note)	Vickers hardness
	Pt	Pd	Cu	P		
20	10	30	40	20	○	490
21	10	40	30	20	◎	480
22	10	50	20	20	○	500
23	10	60	10	20	×	600
24	20	20	40	20	○	510
25	20	30	30	20	◎	470
26	20	40	20	20	◎	460
27	20	50	10	20	×	590
28	30	10	40	20	○	510
29	30	20	30	20	◎	450
30	30	30	20	20	◎	450
31	30	40	10	20	○	500
32	39	2	39	20	×	580
33	40	10	30	20	○	510
34	40	20	20	20	◎	460
35	40	30	10	20	○	500
36	39	39	2	20	×	580
37	2.5	40	37.5	20	×	590
38	5	40	35	20	○	520
39	7.5	40	32.5	20	○	530
40	25	30	25	20	◎	470
41	21	26	21	32	×	590
42	23	29	23	25	○	520
43	27	31	27	15	○	510
44	29	34	29	8	×	600
45	50	10	20	20	○	530
46	50	20	10	20	○	490
47	60	10	10	20	○	520
48	65	5	10	20	○	500
49	70	5	5	20	○	510

◎: Completely vitrified

○: Almost vitrified

×: Crystallization

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As a result of this, an amorphous alloy having a composition within a range recited in claim 2 had a good

[illegible]